

# THE SUCCESS PATH EXAMINATION COUNCIL (SPEC)

Paving the Way to Success

233/3 CHEMISTRY (Practical)

PRE-MOCK EXAMINATIONS 2026

MARCH 2026

Time: 2¼ hours



Paper 3

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## CONFIDENTIAL INSTRUCTIONS TO SCHOOL

The information contained in this paper is to enable the head of school and teacher in charge of **Chemistry** to make adequate preparations for **Pre-Mock 2026 Chemistry practical examination**. NO ONE ELSE should have access to this paper or acquire knowledge of its contents. Great care must be taken to ensure that the information herein does not reach the candidates either directly or indirectly.

The Chemistry teacher is NOT expected to perform the experiments

The apparatus required by each candidate for the Chemistry Pre- mock practical examination are set out on the next page. It is expected that the ordinary apparatus of a chemistry laboratory will be available.

The Chemistry teacher should note that it is his/her responsibility to ensure that each apparatus acquired, for this examination agrees with specifications on the next page.

## REQUIREMENTS FOR CANDIDATES

*In addition to the apparatus and fittings found in the chemistry laboratory, each candidate will require the following;*

1. Exactly 3cm length of **solid A**
2. A bout 80cm<sup>3</sup> of **solution B**
3. About 120cm<sup>3</sup> of **solution D**
4. Means of labeling
5. One 50ml burette
6. One 25ml pipette
7. One pipette filler
8. Two 250ml conical flask
9. One 100ml clean beaker
10. One clean metallic spatula
11. One -10°C to 110°C thermometer
12. Six clean test tubes
13. One test tube holder
14. Two boiling tubes
15. One test tube holder
16. One 10ml measuring cylinder
17. One 250ml volumetric flask
18. A bout 1g of **solid E**
19. About 0.2g of **solid F**
20. About 1g of **solid G**
21. About 2cm<sup>3</sup> of calcium hydroxide solution in stopped test tube.
22. Clean glass rod

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23. About 1g of solid sodium carbonate.
24. About 500cm<sup>3</sup> of distilled water in wash bottle.

### Access to

1. Bunsen burner.
2. 2M aqueous sodium hydroxide with dropper
3. 2M aqueous ammonia supplied with dropper
4. 2M hydrochloric acid
5. Aqueous potassium dichromate (VI) supplied with dropper.
6. Methyl orange indicator supplied with dropper.
7. Phenolphthalein indicator supplied with dropper.

### NOTES

1. Solution **B** is provided by adding 27cm<sup>3</sup> of 98% concentrated sulphuric (VI) acid (sp gr 1.84) to 500cm<sup>3</sup> of distilled water and making the solution 1 litre by adding more distilled water.
2. Solution **C** is provided by dissolving 12g of sodium hydroxide pellets in 500cm<sup>3</sup> of distilled water and then adding more water to make 1 litre of the solution.
3. Calcium hydroxide solution is prepared by adding excess calcium hydroxide powder in cold distilled water and keeping the mixture overnight in a stopped container. The mixture is filtered in the morning of the experiment and supplied in a stopped test tube.
4. Solid **A**, is Magnesium ribbon.
5. Solid **E**, Basic Copper carbonate.
6. Solid **F**; zinc powder.
7. Solid **G**; Oxalic acid

